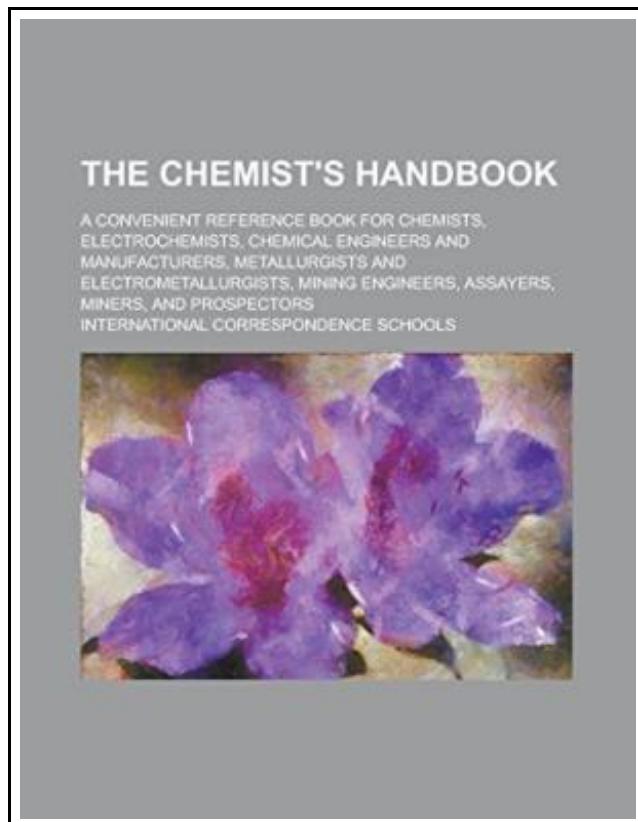


# **The Chemists Handbook A Convenient Reference Book for Chemists, Electrochemists, Chemical Engineers and Manufacturers, Metallurgists and Electrometallurgists, Mining Engineers, Assayers, Miners, and Prospectors**



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## **Reviews**

*Undoubtedly, this is the greatest job by any author. It is actually filled with wisdom and knowledge I am quickly could get a pleasure of reading a written book.*  
**(Kade Ankunding)**

## **THE CHEMISTS HANDBOOK A CONVENIENT REFERENCE BOOK FOR CHEMISTS, ELECTROCHEMISTS, CHEMICAL ENGINEERS AND MANUFACTURERS, METALLURGISTS AND ELECTROMETALLURGISTS, MINING ENGINEERS, ASSAYERS, MINERS, AND PROSPECTORS**

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RareBooksClub. Paperback. Book Condition: New. This item is printed on demand. Paperback. 86 pages. Dimensions: 9.7in. x 7.4in. x 0.2in. This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1922 edition. Excerpt: . . . in 1 l. of distilled water, boil the solution until about one-fifth of it has evaporated, and add sufficient absolute water to bring the volume of the solution up to 1 l. when it is cold. Keep in a glass-stoppered bottle. If this solution becomes green, either in the stock bottle or in the retort when in use, a fresh solution must be made up. **NITROGEN AS NITRITE** It is generally admitted that nitrites found in pond or river water ordinarily come from the putrefaction of nitrogenous organic matter, usually of animal matter. In deep wells, the conditions are somewhat different. These waters may contain nitrites that have been formed by the deoxidation of nitrites in percolating through strata containing mineralreducing agents. The result of this determination should be kept in mind in deciding the quality of a water, and when nitrites are present, the results of the other determinations should be interpreted with greater strictness than when the nitrites are absent. **Determination of Nitrogen as Nitrites.** --Place 50 or 100 c. c. in a Nessler cylinder, add 1 drop of concentrate hydrochloric acid, 1 c. c. of sulphuric-acid solution, and 1 c. c. of naphthylamine. hydrochloride. Mix thoroughly by stirring with a glass rod, cover with a watch glass and set aside for 30 min. Prepare at the same time standard solutions containing known amounts of standard sodium nitrite diluted to 50 or 100 c. c. with absolute water, and the same amount...

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